# metal-organic compounds

 $\gamma = 81.107 (10)^{\circ}$ V = 1006.5 (11) Å<sup>3</sup>

Mo  $K\alpha$  radiation

 $0.10 \times 0.08 \times 0.08$  mm

5579 measured reflections

3848 independent reflections

3001 reflections with  $I > 2\sigma(I)$ 

 $\mu = 1.12 \text{ mm}^{-1}$ 

T = 288 (2) K

 $R_{\rm int} = 0.021$ 

Z = 1

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# Tetra-µ-benzoato-bis[(6-methylquinoline)copper(II)]

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Key indicators: single-crystal X-ray study; T = 288 K; mean  $\sigma$ (C–C) = 0.005 Å; R factor = 0.042; wR factor = 0.101; data-to-parameter ratio = 14.1.

In the title compound,  $[Cu_2(C_7H_5O_2)_4(C_{10}H_9N)_2]$ , the paddlewheel-type dinuclear complex is constructed by four bridging benzoate groups and two terminal 6-methylquinoline ligands. The asymmetric unit contains one-half of the whole molecule, and there is an inversion center at the mid-point of the Cu···Cu bond. The octahedral coordination of each Cu atom, with four O atoms in the equatorial plane, is completed by the N atom of the 6-methylquinoline molecule [Cu-N =2.212 (2) Å] and by another Cu atom [Cu···Cu =2.6939 (13) Å]. The Cu atom lies 0.234 Å out of the plane of the four O atoms. The molecular packing is stabilized by one intramolecular C–H···O as well as C–H··· $\pi$  and  $\pi$ - $\pi$ interactions.

### **Related literature**

For related literature, see: Batten & Robson (1998); Chun *et al.* (2005); Cotton & Walton (1993); Janiak (2003); Lee *et al.* (2008); Mines *et al.* (2002); Pichon *et al.* (2007); Yoo *et al.* (2003).



### **Experimental**

Crystal data

 $\begin{bmatrix} Cu_2(C_7H_5O_2)_4(C_{10}H_9N)_2 \end{bmatrix}$   $M_r = 897.88$ Triclinic,  $P\overline{1}$  a = 10.420 (7) Å b = 10.590 (7) Å c = 10.751 (6) Å  $\alpha = 70.399$  (11)°  $\beta = 64.234$  (10)°

#### Data collection

Bruker SMART CCD area-detector diffractometer Absorption correction: multi-scan (SADABS: Bruker, 1997)  $T_{min} = 0.898, T_{max} = 0.915$ 

#### Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.041$ 272 parameters $wR(F^2) = 0.100$ H-atom parameters constrainedS = 1.04 $\Delta \rho_{max} = 0.31 \text{ e } \text{\AA}^{-3}$ 3848 reflections $\Delta \rho_{min} = -0.33 \text{ e } \text{\AA}^{-3}$ 

#### Table 1

Hydrogen-bond geometry (Å, °).

Cg1 is the centroid of the C22-C27 ring.

$D - H \cdots A$	D-H	$H \cdot \cdot \cdot A$	$D \cdots A$	$D - \mathbf{H} \cdots A$
$C1 - H1 \cdots O11$ $C2 - H2 \cdots Cg1^{i}$	0.93	2.50	3.047 (4)	118
	0.93	2.82	3.734 (3)	168

Symmetry code: (i) x, y, z + 1.

#### Table 2

 $\pi$ - $\pi$  interactions (Å, °).

Cg2 is the centroid of ring C22–C27. The offset is defined as the distance between CgI and the perpendicular projection of CgJ on ring I.

CgI	CgJ	$CgI \cdots CgJ$	Dihedral angle	Interplanar distance	Offset
Cg2	Cg2i	3.967 (4)	0	3.39	2.06

Symmetry code: (i) -x + 2, -y + 2, -z.

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Data collection: *SMART* (Bruker, 1997); cell refinement: *SAINT* (Bruker, 1997); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BX2146).

## References

Batten, S. R. & Robson, R. (1998). Angew. Chem. Int. Ed. 37, 1460-1494.

- Bruker (1997). SMART and SAINT. Bruker AXS Inc., Madison, Wisconsin, USA.
- Chun, H., Dybtsev, D. N., Kim, H. & Kim, K. (2005). Chem. Eur. J. 11, 3521–3529.
- Cotton, F. A. & Walton, R. A. (1993). *Multiple Bonds Between Metal Atoms*, 2nd ed. New York: Oxford University Press.
- Janiak, C. (2003). Dalton Trans. pp. 2781-2804.
- Lee, E. Y., Park, B. K., Kim, C., Kim, S.-J. & Kim, Y. (2008). Acta Cryst. E64, m286.
- Mines, G. A., Tzeng, B.-C., Stevenson, K. J., Li, J. & Hupp, J. T. (2002). Angew. Chem. Int. Ed. 41, 154–157.
- Pichon, A., Fierro, C. M., Nieuwenhuyzen, M. & James, L. (2007). CrystEngComm, 9, 449–451.
- Sheldrick, G. M. (2008). Acta Cryst. A64, 112-122.
- Yoo, S.-K., Ryu, J. Y., Lee, J. Y., Kim, C., Kim, S.-J. & Kim, Y. (2003). Dalton Trans. pp. 1454–1456.